

Please write clearly in block capitals.

Centre number

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Candidate number

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Surname

Forename(s)

Candidate signature

GCSE CHEMISTRY

H

Higher Tier Unit Chemistry C3

Wednesday 14 June 2017

Morning

Time allowed: 1 hour

Materials

For this paper you must have:

- a ruler
- the Chemistry Data Sheet (enclosed).

You may use a calculator.

Instructions

- Use black ink or black ball-point pen.
- Fill in the boxes at the top of this page.
- Answer **all** questions.
- You must answer the questions in the spaces provided. Do not write outside the box around each page or on blank pages.
- Do all rough work in this book. Cross through any work you do not want to be marked.

Information

- The marks for questions are shown in brackets.
- The maximum mark for this paper is 60.
- You are expected to use a calculator where appropriate.
- You are reminded of the need for good English and clear presentation in your answers.
- Question 2(c)(i) should be answered in continuous prose.
In this question you will be marked on your ability to:
 - use good English
 - organise information clearly
 - use specialist vocabulary where appropriate.

Advice

- In all calculations, show clearly how you work out your answer.

For Examiner's Use	
Examiner's Initials	
Question	Mark
1	
2	
3	
4	
5	
6	
TOTAL	



J U N 1 7 C H 3 H P O 1

Answer **all** questions in the spaces provided.

1 This question is about water.

1 (a) Rainwater is soft water.

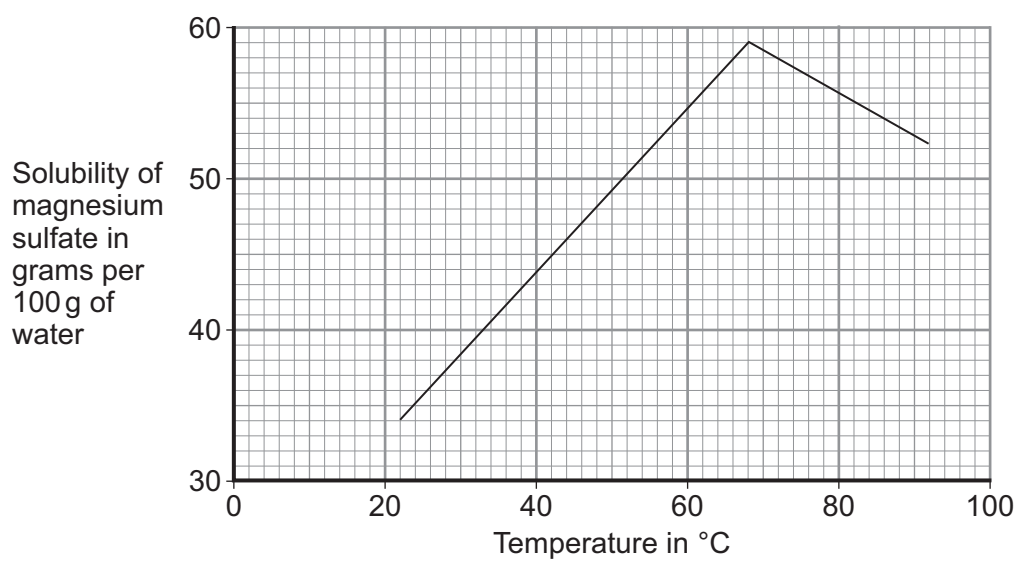
How is hard water formed from rainwater?

[2 marks]

1 (b) A sample of hard water contains magnesium sulfate.

Figure 1 shows the solubility of magnesium sulfate at different temperatures.

Figure 1



What conclusions can be made from **Figure 1**?

Use patterns and values from the graph in your answer.

[3 marks]

1 (c) Give **one** advantage and **one** disadvantage of hard water.

[2 marks]

Advantage _____

Disadvantage _____

1 (d) Describe and explain how hard water is softened using an ion exchange column.

[3 marks]

10

Turn over for the next question

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ANSWER IN THE SPACES PROVIDED**



2 This question is about the combustion of alcohols.

2 (a) What is the structure of methanol?

[1 mark]

Tick (✓) **one** box.

CH₃OH

☐

CH₃CH₂OH

☐

CH₃CH₂CH₂OH

☐

CH₃CH₂CH₂CH₂OH

☐

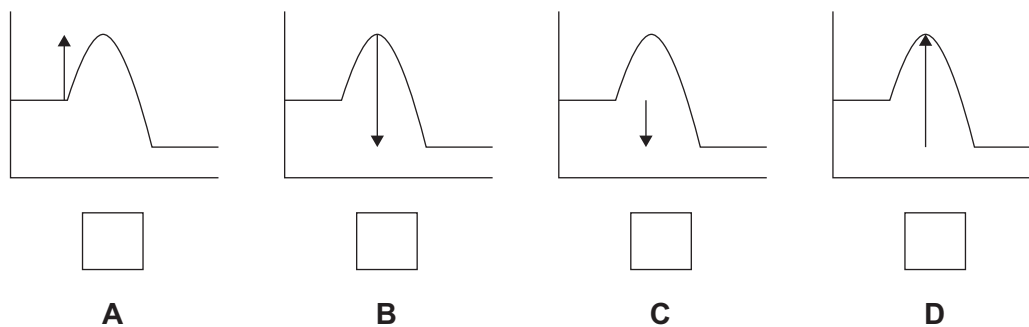
2 (b) **Figure 2** shows four energy level diagrams for the combustion of an alcohol.

Which diagram, **A**, **B**, **C**, or **D**, shows an arrow for the overall energy change?

[1 mark]

Tick (✓) **one** box.

Figure 2



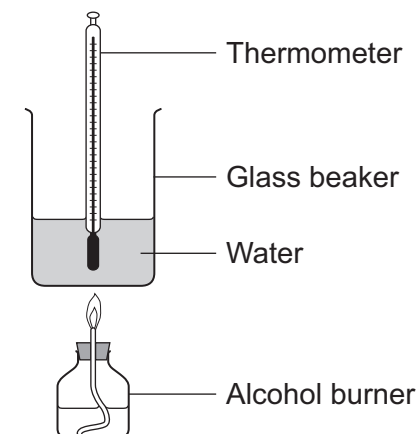
Question 2 continues on the next page

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- 2 (c) **Figure 3** shows apparatus used to measure the energy released when an alcohol is burned.

Figure 3



- 2 (c) (i) **In this question you will be assessed on using good English, organising information clearly and using specialist terms where appropriate.**

Describe how a student could use the apparatus in **Figure 3** to compare the energy released when methanol and ethanol are burned.

You should include any measurements the student would need to make.

Do **not** describe how to do any calculations.

Do **not** describe any improvements to the apparatus.

[6 marks]



Extra space

2 (c) (ii) The student calculated the energy released by the alcohols.

The calculated values were less than the values in a data book.

Explain how the apparatus in **Figure 3** could be improved to obtain more accurate results.

[2 marks]

10

Turn over for the next question

Turn over ►



3 This question is about elements and compounds.

3 (a) In 1869 Mendeleev produced an early version of the periodic table.

Figure 4 shows part of Mendeleev's periodic table.

Figure 4

H						
Li	Be	B	C	N	O	F
Na	Mg	Al	Si	P	S	Cl
K	Ca		Ti	V	Cr	Mn

3 (a) (i) Why did Mendeleev leave gaps in his periodic table?

[1 mark]

3 (a) (ii) Give **one** reason why the elements Ti, V, Cr and Mn should **not** be where Mendeleev placed them.

[1 mark]



3 (b) **Figure 5** shows an outline of the modern periodic table.

Figure 5

[illegible]

3 (b) (i) Why, in terms of electrons, is bromine in Group 7?

[1 mark]

3 (b) (ii) Bromine reacts with sodium iodide.

The word equation for the reaction is:



This reaction shows that bromine is more reactive than iodine.

Explain why, in terms of electrons.

[3 marks]

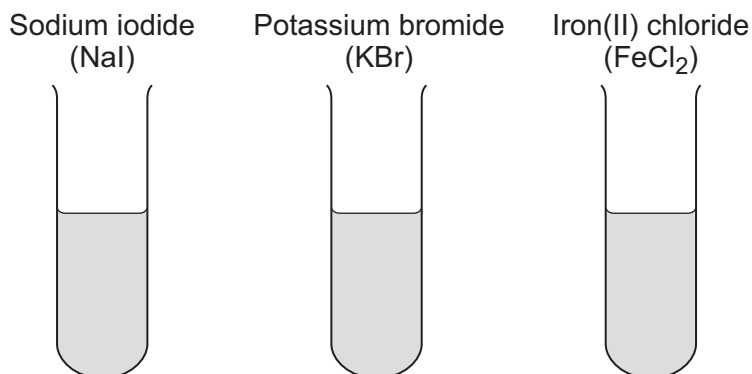
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- 3 (c)** **Figure 6** shows test tubes containing solutions of sodium iodide, potassium bromide and iron(II) chloride.

Figure 6



- 3 (c) (i)** A student tested each solution with silver nitrate in the presence of dilute nitric acid.

Write the result for sodium iodide solution in **Table 1**.

[1 mark]

Table 1

	Sodium iodide solution	Potassium bromide solution	Iron(II) chloride solution
Result of adding silver nitrate		Cream precipitate	White precipitate

- 3 (c) (ii)** The student tested new samples of each solution with sodium hydroxide solution.

Write the result for iron(II) chloride solution in **Table 2**.

[1 mark]

Table 2

	Sodium iodide solution	Potassium bromide solution	Iron(II) chloride solution
Result of adding sodium hydroxide solution	No precipitate	No precipitate	



- 3 (d)** A flame test is done on a mixture of sodium iodide and potassium bromide.

Why would a flame test **not** show the presence of both sodium ions and potassium ions in the mixture?

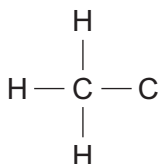
[1 mark]

- 3 (e)** Carboxylic acids are compounds containing carbon.

- 3 (e) (i)** Complete **Figure 7** to show the displayed structure of ethanoic acid.

[1 mark]

Figure 7



- 3 (e) (ii)** Explain why ethanoic acid has a higher pH value than hydrochloric acid of the same concentration.

[2 marks]

12

Turn over for the next question

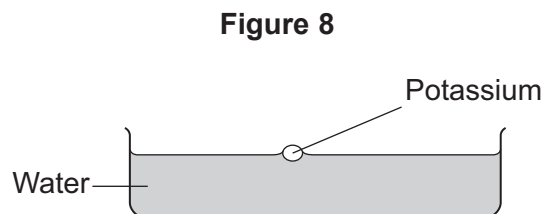
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4 This question is about potassium and its compounds.

4 (a) (i) Potassium reacts with water to produce potassium hydroxide solution and a gas.

Figure 8 shows the apparatus used for the reaction.



Complete and balance the equation for the reaction.

[2 marks]



4 (a) (ii) Potassium and lithium are in Group 1 of the periodic table.

Give **two** differences you would see between the reactions of potassium and lithium with water.

[2 marks]



- 4 (b)** Describe how a student could do titrations to find the mean volume of potassium hydroxide solution which would neutralise 25.00 cm³ of nitric acid.

[5 marks]

- 4 (c)** The student found that 26.25 cm³ of potassium hydroxide solution with a concentration of 0.20 moles per dm³ neutralises 25.00 cm³ of nitric acid.

The equation for the reaction is:



Calculate the concentration of the nitric acid.

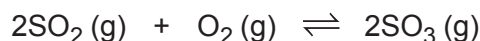
[3 marks]

Concentration of nitric acid = _____ moles per dm³



5 This question is about sulfuric acid and sulfates.

5 (a) One reaction in the industrial production of sulfuric acid is:



The forward reaction is exothermic.

5 (a) (i) Explain why a moderately high temperature, instead of a low temperature, is used for this reaction.

Use the information above and your knowledge.

[3 marks]

5 (a) (ii) Explain why you would expect this reaction to be carried out at high pressure.

[2 marks]

5 (a) (iii) This reaction is carried out at atmospheric pressure.

Suggest **one** advantage, other than cost, of using atmospheric pressure and **not** a high pressure for this reaction.

[1 mark]



5 (b) Sulfuric acid is used to produce sulfates.

Describe how sodium hydroxide solution is used to distinguish between a solution of magnesium sulfate and a solution of aluminium sulfate.

[2 marks]

8

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6 This question is about ammonia and a compound produced from ammonia.

6 (a) Ammonia is produced by the Haber process.

6 (a) (i) The reactants in the Haber process are nitrogen and hydrogen.

Give the source of the hydrogen used for the Haber process.

[1 mark]

6 (a) (ii) The Haber process produces a mixture of the gases ammonia, hydrogen and nitrogen.

Table 3 shows boiling points for ammonia, hydrogen and nitrogen.

Table 3

	Boiling point in °C
Ammonia	–33
Hydrogen	–260
Nitrogen	–196

Explain how to separate ammonia from the mixture of gases.

Use the information in **Table 3** to help you.

[2 marks]

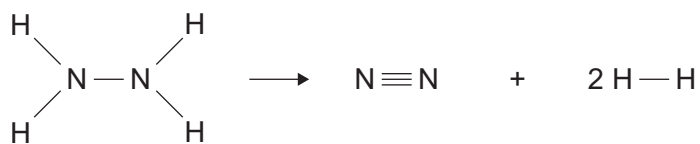
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6 (b) Ammonia is used to produce hydrazine (N_2H_4).

6 (b) (i) The equation for the decomposition of hydrazine is:



Bond energies are given in **Table 4**.

Table 4

Bond	Bond energy in kJ
$\text{N} - \text{N}$	160
$\text{N} - \text{H}$	391
$\text{N} \equiv \text{N}$	941
$\text{H} - \text{H}$	432

Calculate the energy change for the decomposition of hydrazine.

[3 marks]

Energy change = _____ kJ



6 (b) (ii) The decomposition of hydrazine is exothermic.

Explain why, in terms of bond making and bond breaking.

[2 marks]

8

END OF QUESTIONS



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